

GCSE Chemistry B (Twenty First Century Science)
J258/01 Breadth in Chemistry (Foundation Tier)

Question Set 18

Multiple Choice Questions

1

A chemist makes a teaspoon out of gallium metal.

Gallium looks like aluminium. Gallium melts at 30 °C and aluminium melts at 660 °C.



(a) Tea is made with boiling water.

What would happen if a gallium spoon is used to stir hot tea?

Explain your answer.

[2]

(b) When Mendeleev made his Periodic Table, he left a gap below aluminium.

Later gallium was discovered and put into this gap.

Give one reason why gallium fitted into this gap.

Tick (✓) **one** box.

It has a similar melting point to aluminium.

It looks the same as aluminium.

It has similar reactions to aluminium.

There was nowhere else in the table to put it.

[1]

(c) When gallium reacts it loses three electrons.

Which ion is formed?

Put a **ring** around the correct answer.

Ga Ga⁺ Ga²⁺ Ga³⁺ Ga⁻ Ga²⁻ Ga³⁻

[1]

(d) Ionic compounds have high melting points.

Which two statements explain this?

Tick (✓) **two** boxes.

There are strong attractions between the ions.

Shared electron bonds are broken.

A lot of energy is needed to separate the ions.

Positive ions attract other positive ions.

Ionic compounds conduct electricity.

[2]

(e) Mendeleev put potassium and sodium in the same group because they both react with water.

Ali's teacher puts a piece of sodium into water. The teacher then puts a piece of potassium into water.

Give **two** ways Ali could tell potassium is more reactive than sodium.

[2]

Total Marks for Question Set 18: 8

Resource Materials

Question Set No: 18

The Periodic Table of the Elements

(1)	(2)											(3)	(4)	(5)	(6)	(7)	(0)	
1 H hydrogen 1.0																		2 He helium 4.0
3 Li lithium 6.9	4 Be beryllium 9.0											5 B boron 10.8	6 C carbon 12.0	7 N nitrogen 14.0	8 O oxygen 16.0	9 F fluorine 19.0	10 Ne neon 20.2	
11 Na sodium 23.0	12 Mg magnesium 24.3											13 Al aluminium 27.0	14 Si silicon 28.1	15 P phosphorus 31.0	16 S sulfur 32.1	17 Cl chlorine 35.5	18 Ar argon 39.9	
19 K potassium 39.1	20 Ca calcium 40.1	21 Sc scandium 45.0	22 Ti titanium 47.9	23 V vanadium 50.9	24 Cr chromium 52.0	25 Mn manganese 54.9	26 Fe iron 55.8	27 Co cobalt 58.9	28 Ni nickel 58.7	29 Cu copper 63.5	30 Zn zinc 65.4	31 Ga gallium 69.7	32 Ge germanium 72.6	33 As arsenic 74.9	34 Se selenium 79.0	35 Br bromine 79.9	36 Kr krypton 83.8	
37 Rb rubidium 85.5	38 Sr strontium 87.6	39 Y yttrium 88.9	40 Zr zirconium 91.2	41 Nb niobium 92.9	42 Mo molybdenum 95.9	43 Tc technetium	44 Ru ruthenium 101.1	45 Rh rhodium 102.9	46 Pd palladium 106.4	47 Ag silver 107.9	48 Cd cadmium 112.4	49 In indium 114.8	50 Sn tin 118.7	51 Sb antimony 121.8	52 Te tellurium 127.6	53 I iodine 126.9	54 Xe xenon 131.3	
55 Cs caesium 132.9	56 Ba barium 137.3	57-71 lanthanoids	72 Hf hafnium 178.5	73 Ta tantalum 180.9	74 W tungsten 183.8	75 Re rhenium 186.2	76 Os osmium 190.2	77 Ir iridium 192.2	78 Pt platinum 195.1	79 Au gold 197.0	80 Hg mercury 200.6	81 Tl thallium 204.4	82 Pb lead 207.2	83 Bi bismuth 209.0	84 Po polonium	85 At astatine	86 Rn radon	
87 Fr francium	88 Ra radium	89-103 actinoids	104 Rf rutherfordium	105 Db dubnium	106 Sg seaborgium	107 Bh bohrium	108 Hs hassium	109 Mt meitnerium	110 Ds darmstadtium	111 Rg roentgenium	112 Cn copernicium		114 Fl flerovium		116 Lv livermorium			

Key atomic number Symbol name relative atomic mass
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